

CHEM 121: General Chemistry I

Spring 2019 – Section 501 – CRN 36461

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Office Hours: Monday 1:00 p.m. – 4:00 p.m.
Tuesday 3:00 p.m. – 4:00 p.m.
Wednesday 3:00 p.m. – 4:00 p.m. and other times by appointment

Meeting Times: Lecture: Monday & Wednesday 9:00 – 10:15 am, VAAS 131
Laboratory: Monday 10:30 am – 1:15 pm, VAAS 128

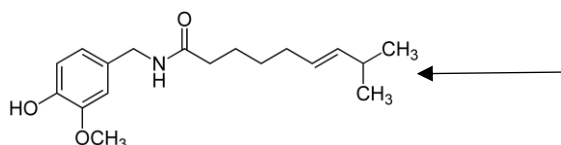
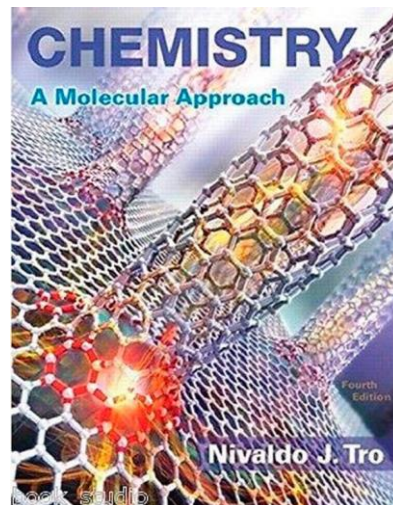
Course Description: The Study of stuff, and what it does (1st of a 2-course sequence)

Course Description: Introduction to the chemical and physical behavior of matter. Meets New Mexico Lower Division General Education Common Core Curriculum Area III: Science (NMCCN 1214). Prerequisite: MATH 121 or MATH 123 or MATH 150 or MATH 153 or MATH 162 or MATH 163 or MATH 180 or MATH 181 or MATH 264 or ACT Math =>25 or SAT Math Section =>590. Pre- or corequisite: 123L. {Summer, Fall, Spring}

Guess which one is the instructor's, and guess which one is has gone through various committees and perhaps a lawyer or two?

Periodic Table of the Elements

1	Periodic Table of the Elements																18
H Hydrogen 1.008																	He Helium 4.003
3	4											5	6	7	8	9	10
Li Lithium 6.941	Be Beryllium 9.012											B Boron 10.811	C Carbon 12.011	N Nitrogen 14.007	O Oxygen 15.999	F Fluorine 18.998	Ne Neon 20.180
11	12											13	14	15	16	17	18
Na Sodium 22.990	Mg Magnesium 24.305											Al Aluminum 26.982	Si Silicon 28.086	P Phosphorus 30.974	S Sulfur 32.06	Cl Chlorine 35.453	Ar Argon 39.948
19	20	21	22	23	24	25	26	27	28	29	30	31	32	33	34	35	36
K Potassium 39.098	Ca Calcium 40.078	Sc Scandium 44.956	Ti Titanium 47.88	V Vanadium 50.942	Cr Chromium 51.996	Mn Manganese 54.938	Fe Iron 55.845	Co Cobalt 58.933	Ni Nickel 58.693	Cu Copper 63.546	Zn Zinc 65.38	Ga Gallium 69.723	Ge Germanium 72.63	As Arsenic 74.922	Se Selenium 78.971	Br Bromine 79.904	Kr Krypton 83.798
37	38	39	40	41	42	43	44	45	46	47	48	49	50	51	52	53	54
Rb Rubidium 84.96	Sr Strontium 87.62	Y Yttrium 88.906	Zr Zirconium 91.224	Nb Niobium 92.906	Mo Molybdenum 95.94	Tc Technetium 98.906	Ru Ruthenium 101.07	Rh Rhodium 102.905	Pd Palladium 106.42	Ag Silver 107.868	Cd Cadmium 112.414	In Indium 114.818	Sn Tin 118.710	Sb Antimony 121.757	Te Tellurium 127.6	I Iodine 126.905	Xe Xenon 131.29
55	56	57-71	72	73	74	75	76	77	78	79	80	81	82	83	84	85	86
Cs Cesium 132.905	Ba Barium 137.327	Lanthanides	Hf Hafnium 178.49	Ta Tantalum 180.948	W Tungsten 183.84	Re Rhenium 186.207	Os Osmium 190.23	Ir Iridium 192.225	Pt Platinum 195.084	Au Gold 196.967	Hg Mercury 200.592	Tl Thallium 204.38	Pb Lead 207.2	Bi Bismuth 208.980	Po Polonium [209]	At Astatine [210]	Rn Radon [222]
87	88	89-103	104	105	106	107	108	109	110	111	112	113	114	115	116	117	118
Fr Francium [223]	Ra Radium [226]	Actinides	Rf Rutherfordium [261]	Db Dubnium [262]	Sg Seaborgium [266]	Bh Bohrium [264]	Hs Hassium [277]	Mt Meitnerium [276]	Ds Darmstadtium [285]	Rg Roentgenium [282]	Cn Copernicium [285]	Uut Ununtrium [288]	F1 Flerovium [288]	Uup Ununpentium [294]	Lv Livermorium [293]	Uus Ununseptium [294]	Uuo Ununoctium [294]
57	58	59	60	61	62	63	64	65	66	67	68	69	70	71			71
La Lanthanum 138.905	Ce Cerium 140.116	Pr Praseodymium 140.908	Nd Neodymium 144.242	Pm Promethium [145]	Sm Samarium 150.36	Eu Europium 151.964	Gd Gadolinium 157.25	Tb Terbium 158.925	Dy Dysprosium 162.50	Ho Holmium 164.930	Er Erbium 167.259	Tm Thulium 168.934	Yb Ytterbium 173.054	Lu Lutetium 174.967			Lu Lutetium [175]
89	90	91	92	93	94	95	96	97	98	99	100	101	102	103			103
Ac Actinium [227]	Th Thorium 232.038	Pa Protactinium 231.036	U Uranium 238.029	Np Neptunium 237.048	Pu Plutonium 244.064	Am Americium [243]	Cm Curium [247]	Bk Berkelium [247]	Cf Californium [251]	Es Einsteinium [252]	Fm Fermium [257]	Md Mendelevium [258]	No Nobelium [259]	Lr Lawrencium [262]			Lr Lawrencium [262]



What is this molecule? Tell me (email) for some extra credit!

WHAT YOU'LL NEED (Required Resources)

- Chemistry: A Molecular Approach
- Mastering Chemistry Access Code
- Calculator (non-graphing) with log/antilog and exponential functions
- Internet Access: *Blackboard Learn* and *UNM email address* **must be checked regularly (daily)**

WHAT IF YOU NEED HELP? (UNM-Valencia Resources)

- **Instructor:** Office hours, STEM Center Hours, email
- **STEM Center:** Tutors*, molecular modelling kits, Laptops, textbooks

* Reminder: when using tutors, it is the **students'** responsibility to make sure they understand well enough to complete the problems on **their own**.

WHAT WILL EACH CLASS BE LIKE?

- **Quiz:** covering material recently covered and any assigned preparation (reading, video, etc)
- **Course Business**
- **Group Activity:** collaborative exercise to help you master that day's topic
- **Reflection:** an opportunity to put the day's lesson into larger perspective, and formulate/ask questions

WHAT WILL MY ROUTINE BE LIKE?

- **Before Class:** complete any preparatory assignment (reading, video, etc.)
- **During Class:** work with your group to master concepts. The more you put in, the more you'll get out
- **After Class:** work on homework assignment relevant to that day's topic (review notes, **WORK ON PROBLEMS**, think of questions for office hour visits, etc.
- **Repeat 28 times!:**

WHAT YOU'LL FIND USEFUL (Recommended Resources)

- 3-ring binder for lecture notes, handouts, group activities
- Periodic table (on paper)
- Calculator (non-graphing) with log/antilog and exponential functions
- Mastering Chemistry notebook: keep track of problem solving, identify patterns, record areas of difficulty

HOW IS YOUR GRADE DETERMINED?

(Exams, Quizzes, Homework, and the Like)

	How Many	Weight
Class Points	1	10 %
Quizzes	15*	10 %
Homework	10*	15 %
Exams	4	50 %
Final Exam	1	15 %
Total		100 %

* Approximate values

** Each equally weighted, 12.5 % each

WHAT DO I NEED FOR AN A?

(What's the grading scale?)

Earn This %	Get This Grade
98	A+
92	A
90	A-
88	B+
83	B
80	B-
78	C+
73	C
69	C-
67	D+
62	D
60	D-
55	F+
0	F

WHEN WE LEARN THIS STUFF?

(Schedule is approximate and subject to change by the instructor)

Class	Date	Topics/Events
1	Mon 14 Jan	Nuclear Atom GA (2.5 – 2.6)
2	Wed 16 Jan	Dimensional Analysis GA (1.6 – 1.8)
3	Mon 21 Jan	Martin Luther King Day – No meeting
4	Wed 23 Jan	Dalton's Atomic Theory (2.1 – 2.5)
5	Mon 28 Jan	Periodic Table, Average Atomic Mass GA (2.7 – 2.8)
6	Wed 30 Jan	Molar Mass (Counting by Weighing GA) (2.8)
7	Mon 04 Feb	Chemical Bonding, Formulas and Naming (3.1 – 3.7)
8	Wed 06 Feb	Exam 1 (Chapters 1 & 2)
9	Mon 11 Feb	Molar Mass, Balanced Chemical Equations (3.8 – 3.12)
10	Wed 13 Feb	Stoichiometry, L.R., % Yield, (4.1 – 4.2)
11	Mon 18 Feb	Solution Stoichiometry. Aqueous Solutions, Molarity (4.3 – 4.6)
12	Wed 20 Feb	Aqueous Reactions, Net Ionic Equations (4.7 – 4.9)
13	Mon 25 Feb	Ideal Gas Equation (5.1 – 5.4)
14	Wed 27 Feb	Exam 2 (Chapters 3 – 4)
15	Mon 04 Mar	Gas Mixtures, Gas Stoichiometry (5.6 – 5.7)
16	Wed 06 Mar	Kinetic Molecular Theory, Real Gases, Thermodynamics Intro (5.8, 5.10, 6.1 – 6.3)
	Mon 11 Mar	Spring Break (no meeting) Online Thermodynamics Activity
	Wed 13 Mar	Spring Break (no meeting)
17	Mon 18 Mar	Thermochemistry and Calorimetry (6.3 – 6.7)
18	Wed 20 Mar	Hess' Law and Reaction Enthalpies (6.8 – 6.9)
19	Mon 25 Mar	Atomic Orbitals and Shapes (7.5 – 7.6)
20	Wed 27 Mar	Electronic Configurations and Periodic Table (8.1 – 8.5)
21	Mon 01 Apr	Exam 3 (Chapters 5 – 7)
22	Wed 03 Apr	Periodic Trends (8.6 – 8.9)
23	Mon 08 Apr	Lewis Dot Structures (9.1 – 9.5)
24	Wed 10 Apr	Bond Polarity, Dipoles, Bond Characteristics (9.6 – 9.11)
25	Mon 15 Apr	VSEPR Theory (10.1 – 10.5)
26	Wed 17 Apr	Hybridization (10.6 – 10.7)
27	Mon 22 Apr	Bonding Review
28	Wed 24 Apr	Exam 4 (Chapters 8 – 10)
29	Mon 29 Apr	Molecular Orbital Theory (10.8)
30	Wed 1 May	Molecular Orbital Theory (10.8)
	Mon 6 May	Final Exam (9:00 – 11:00 a.m.)

Class Policies and Important Dates

- | | |
|---|--|
| <ul style="list-style-type: none"> • Be There Attendance in lecture and lab/recitation is mandatory. Students are expected to attend all meetings of the classes in which they are enrolled. <ul style="list-style-type: none"> ○ A student with excessive absences may be dropped from a course by the instructor with a grade of WP or WF or the student may receive a grade of F at the end of the semester. ○ I will exercise my discretion without notice to drop any student who: <ul style="list-style-type: none"> ▪ misses the first two meetings; ▪ has not completed any assignments in BB Learn and/or Mastering Chemistry by the end of the 2nd week; ▪ after 2 consecutive unexcused absences; or after 4 total absences. ○ Excused absences must be authorized. | <ul style="list-style-type: none"> • Be on time. Lectures and labs/recitations will begin promptly. After 10 minutes, a student will be counted absent. Late arrival or early departure is unacceptable. Absences due to illness or any mitigating circumstance are unavoidable but must be documented or approved in advance. If you must miss a lecture or lab, email me ASAP in order to get your absence excused and discuss when you will turn in or make up any allowable assignments. Students are responsible for all assignments regardless of attendance. • Your job begins when class ends: Electronic homework will be assigned regularly. Your answers are to be submitted and scored on Mastering Chemistry. Late homework will not be accepted. |
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Important Dates & Holidays

Fri 25 Jan 2019	Last day to register, ADD sections, and change credit hours Enrollment cancellation for non-payment
Mon 21 Jan 2019	University Holiday – Martin Luther King Day
Fri 01 Feb 2019	Last Day to DROP without “W” grade and 100% tuition refund on LoboWEB
Fri 08 Feb 2019	Last Day to CHANGE grade option
Sun 10 Mar 2019	Spring Break (through Sat 16 Mar 2019)
Fri 12 April 2018	Last Day to withdraw WITHOUT Dean’s Permission
Fri 03 May 2018	Last Day to withdraw WITH Dean’s Permission
Mon 06 May 2019	Final Exam (for this section)

Course Learning Objectives

Unit Level Learning Objectives: Exam 1 (Chapters 1-2)

At the end of most learning objectives, there is a reference to a sample problem. These references are the same for both the 3rd and 4th editions of the textbook. The following symbols are used for these references:

CC = Conceptual connection problem within the chapter (answers at the end of the chapter)

Ex = Example within the chapter

EoC = End of chapter problems (answers in Appendix III)

By the end of the chapter, students will be able to...

Ch 1: Matter, Measurement, and Problem Solving

1. Define matter and classify a given substance by physical state.
2. Classify changes in matter as physical or chemical. (Ex 1.1 p 10)
3. Use the appropriate SI units and metric prefixes to express numbers in scientific notation. (Ex 1.4 p 21, Ex 1.5 p 23)
4. Use the concept of density in quantitative and qualitative problems involving masses and volumes. (Ex 1.8 p 29, Ex 1.10 p 31)
5. Report the result of any measurement to the appropriate number of significant figures. (Ex 1.6 p 24)
6. Express the result of any set of simple mathematical operations on measurements to the appropriate number of significant figures. (Ex 1.4 p 21, Ex 1.5 p 23, Ex 1.6 p 24)
7. Analyze a set of measurements for precision and or accuracy. (Ex 1.7-1.8 p 29)
8. Convert between units and prefixed units using dimensional analysis and develop a systematic approach to solving problems involving unit conversion and equations, including the conversion between the three commonly used temperature scales. (Ex 9 p 30, Ex 1.10 p 31)

Ch 2: Atoms and Elements

1. Use the laws of conservation of mass, definite proportions, and multiple proportions to justify Dalton's atomic theory. (Ex 2.1 p 49, Ex 2.2 p 50, CC 2.2 p 50)
2. Justify the nuclear model of the atom with reference to Rutherford, Thompson's,

Millikan's experiments, and the scientific method. (CC 2.3 p 53).

3. Identify a set of isotopes from information on the composition of the nucleus. Use atomic notation to write the symbol of any isotope. (Ex 2.3 p 59, CC 2.4 p 59)
4. Identify an element or ion based on the composition of the nucleus and number of electrons. (CC 2.5 p 61, Ex 2.4 p 65)
5. Use the periodic table to classify an element as being a metal (forms cations), nonmetal (forms anions).
6. Identify main group elements and transition elements. Also identify the following groups: alkali metals, alkaline earth metals and halogens and recall the ions commonly formed by elements in these groups.
7. Define the mole and calculate and use average atomic masses to convert between mass, moles and numbers of atoms. (Ex 2.6 p 71, Ex 2.7 p 72, Ex 2.8 p 73, Ex 2.9 p 74)

Unit Level Learning Outcomes: Exam 2 (Chapters 3-4)

At the end of most learning objectives, there is a reference to a sample problem. These references are the same for both the 3rd and 4th editions of the textbook.

The following symbols are used for these references:

CC = Conceptual connection problem within the chapter (answers at the end of the chapter)

Ex = Example within the chapter

EoC = End of chapter problems (answers in Appendix III)

By the end of the chapter, students will be able to...

Ch 3: Molecules, Compounds, Chemical Equations

1. Describe the two different forms of bonding that connect atoms - IONIC or COVALENT. Use the periodic table to determine whether a species is molecular or ionic based on chemical formula. (EoC 29 p 130)
2. Determine formulas of ionic compounds, including the use of polyatomic ions, and molecules from their systematic names. (EoC 33 & 35 p 131, Ex 3.2 p 95)
3. Name molecular and ionic compounds using their systematic names. (EoC 37, 41, 47, 49 p131)

- Determine and use molar mass to convert between mass, moles, and numbers of molecules and atoms in molecules. (Ex 3.13 p 108)
- Write and balance chemical equations to describe reactions. (Ex 3.22, 3.23, 3.24 p 120-122)

Ch 4: Chemical Quantities and Aqueous Reactions

- Define molarity and perform calculations involving the composition of solutions, including dilution calculations. (Ex 4.1 p 143, Ex 4.2 p 144, Ex 4.5 p 153, Ex 4.7 p 156)
- Define and give examples of strong electrolytes, weak electrolytes, and non-electrolytes. Draw molecular level pictures of each type of electrolyte to illustrate the relative degree of ionization in each.
- Determine the products of a given precipitation reaction by considering the species present in solution and using a solubility table. (Ex 4.10 & 4.11 p 165)
- Represent precipitation, acid-base, and gas evolution reactions in solution by molecular, complete ionic, and net ionic equations. (Ex 4.12 p 168, Ex 4.13 p 171)
- Perform stoichiometric calculations involving precipitation reactions or acid-base neutralization reactions, including those involving limiting reagent. (Ex 4.14 p 173)
- Define oxidation and reduction in terms of electron loss and gain. (Ex 4.17 p 179)
- Assign oxidation states to simple ionic compounds and use oxidation state changes to identify redox reactions, oxidizing and reducing agents. (Ex 4.16 p 178, CC 4.8 p 179)
- Write balanced equations for combustion reactions, precipitation, and acid-base reactions. (Ex 4.18 p 180, Ex 4.19 p 182)

Unit Level Learning Outcomes: Exam 3 (Chapters 5-7)

At the end of most learning objectives, there is a reference to a sample problem. These references are the same for both the 3rd and 4th editions of the textbook. The following symbols are used for these references:

CC = Conceptual connection problem within the chapter (answers at the end of the chapter)

Ex = Example within the chapter

EOC = End of chapter problems (answers in Appendix III)

By the end of the chapter, students will be able to...

Ch 5: Gases

- Recall and use the gas laws (Boyle, Charles and Avogadro) to calculate properties of an ideal gas under changing conditions. (Ex 5.2 p 202, Ex 5.3 p 205, CC 5.1 p 205)
- Recall and use the ideal gas law, $PV = nRT$ to calculate P, V, n or T given three of the four parameters. (Ex 5.5 p 208, Ex 5.6 p 209)
- Recall and use the molar volume for an ideal gas 22.42 L at STP (recall that STP is 0 °C (273K) and 1atm). (CC 5.2 p 210, CC 5.3 p 211, Ex 5.7 p 213)
- Recall and apply Dalton's Law of Partial Pressures to calculate properties relating to mixtures of gases. Use and calculate mole fractions. (CC 5.4 p 216, Ex 5.9 p 216, Ex 5.10 p 218)
- Apply the ideal gas law to find number of moles from P, V and T conditions, and use this information in stoichiometric calculations. (Ex 5.12 p 221, Ex 5.13 p 223)
- Recall the three assumptions of Kinetic Molecular Theory and identify situations in which these assumptions fail.

Ch 6: Thermochemistry

- Define potential energy, kinetic energy and work.
- State the first law of thermodynamics.
- Distinguish between heat and temperature. (CC 6.2 p 257)
- Identify chemical bonds as the source of chemical potential energy.
- Define energy flow INTO a system as a positive quantity, and energy flow OUT of a system as a negative quantity for the system. Apply the terms 'endothermic' and 'exothermic' to describe the flow of heat between a reaction and its surroundings. Relate these terms to the relative chemical potential energy of reactant and products. (Table 6.3 p 256)
- Define and use specific and molar heat capacities to calculate temperature changes when heat is applied or removed. (CC 6.3 p 260, Ex 6.3 p 261)
- Apply stoichiometry to determine enthalpy changes associated with reactions of particular masses of reactants or to form particular amounts of products. (Ex 6.7 p 270, CC 6.5 p 267)
- Use specific or molar heat capacities to calculate the enthalpy of a reaction in a calorimeter (constant pressure or constant volume). (Ex 6.5 p 266, Ex 6.8 p 271)
- Use the properties of enthalpy to calculate ΔH for a chemical reaction using Hess's Law. (Ex 6.9 p 274)
- Look up standard enthalpies of formation for any substance and apply these to calculate ΔH for a reaction. (Ex 6.10 p 276)

By the end of the chapter, students will be able to...

Ch 7: Electronic Structure of Atoms

- Use the emission spectrum of hydrogen in the visible region to explain how this line spectrum

- supports a quantized model of energy levels in hydrogen. (Ex 7.7 p 322)
2. Describe the Bohr model of the hydrogen atom in terms of quantized circular orbits.
 3. Use quantum numbers n , l , and m_l to describe orbitals. Recall and use the relationships between n , l and m_l to determine if any orbital is an allowed one, what type of orbital it is (s , p , d or f orbital), and how many orbitals there are in each l level. (CC 7.4-7.5 p 318, Ex 7.5-7.6 p 320)
 4. Sketch the shapes of orbitals designated by s , p , and d . (Figure 7.28 p 327)

Unit Level Learning Outcomes: Exam 4 (Chapters 8-10)

Ch 8: Periodic Properties

1. Write electron configurations and orbital diagrams for ground state atoms by applying the Pauli exclusion principle, Hund's rule, the Aufbau principle, and the position of the atom in the Periodic Table. Identify atoms based on electron configurations and orbital diagrams. (Fig 8.5 p 343, Ex 8.1-8.2 p 346)
2. Identify the principle quantum number and the number of valence electrons for an atom or ion and use this information to predict the relative reactivity, size, magnetism, and ionization energy of the atom or ion. (Ex 8.3 p 347, Ex 8.4 p 350)
3. Understand the concept of effective nuclear charge and how it affects atomic size. (Ex 8.5 p 356, Fig 8.12 p 359)

By the end of the chapter, students will be able to...

Ch 9: Lewis Model of Bonding

1. Describe covalent and ionic bonding with respect to orbitals. (CC 9.1 p 386, Ex 9.3 p 400)
2. Use Lewis structures to represent the valence electrons of molecules and determine bond order and placement of non-bonding electrons. (Ex 9.1 p 388, Ex 9.4-9.5 p 401, Ex 9.6 p 402)
3. Use formal charge considerations to determine the lowest energy resonance structure for a molecule. (Ex 9.7 p 404, Ex 9.8 p 406)

4. Use trends in electronegativity to determine bond polarity. Predict the relative polarity of covalent bonds. (CC 9.4 p 398)
5. Predict relative bond energies and bond lengths in related molecules. (CC 9.8 p 414, Ex 9.11 p 414)

Ch 10: VSEPR and Molecular Orbital Theory

1. Predict the shape of any given molecule by writing the Lewis structure and applying VSEPR to assign the positions of the bonding and non-bonding electrons pairs. (CC 10.1 p 429, CC 10.2 p 431, Ex 10.1 p 432)
2. Compare bond angles in the series methane, ammonia and water to demonstrate that lone pairs repel more than bonded pairs of electrons. (Ex 10.2-10.3 p 438)
3. Draw dipole moments for bonds in molecules, and use these to predict whether a molecule will have a net dipole moment. (Ex 10.5 p 443)
4. Explain what hybridization is and why we invoke it in Valence Bond theory to describe bonding in covalent compounds.
5. Determine the appropriate hybridization of any atom in a molecule using the Lewis structure and the number of electron groups in it (2 to 6 groups). (CC 10.7 p 450)
6. Show how orbitals overlap to form new orbitals with sigma or pi symmetry. Explain why sigma overlap is greater than pi overlap and describe the implications for bond strength. (CC 10.8 p 454)
7. Analyze a given organic 'skeleton' structure to determine geometry of any given atom and the number of sigma bonds and pi-bonds in the structure. (Ex 10.6-10.7 p 459, Ex 10.8 p 460)
8. Draw molecular orbital diagrams for homonuclear diatomics from hydrogen to fluorine and their anion and cation forms. Use MO diagrams to predict bond order, relative bond lengths and strengths, and paramagnetism. (Ex 10.9 p 464, Fig 10.5 p 468)

Other Things That Aren't Chemistry, But Are Still Important (University Policies)

Equal Access Services

If you have a documented disability or psychological/medical condition that may affect your performance in this class, please register with Equal Access Services as soon as possible so I can provide your accommodations in a timely manner. EAS can provide a quiet place to take exams, additional time, and additional services if there is a documented need. For more information, please see their website at



Equal Access Services

<https://valencia.unm.edu/students/advisement-and-counseling/equal-access-services.html>, or scan the QR code above:

Academic Integrity

Having academic integrity is paramount to your success in any class. Plagiarism or cheating is not tolerated. Any instance of this will result in a grade of zero for that assignment.



Academic Integrity Policy

Here is the link to the UNM Academic Dishonesty Policy:

<https://policy.unm.edu/regents-policies/section-4/4-8.html>, or scan the QR code above:

The policy states:

Each student is expected "to maintain the highest standards of honesty and integrity in academic and professional matters. The University reserves the right to take disciplinary action, up to and including dismissal, against any student who is found guilty of academic dishonesty or who otherwise fails to meet the expected standards. Any student judged to have engaged in

academic dishonesty in course work may receive a reduced or failing grade for the work in question and/or for the course.

Academic Dishonesty is defined as:

"Academic dishonesty" includes, but is not limited to, dishonesty in quizzes, tests, or assignments; claiming credit for work not done or done by others; hindering the academic work of other students; misrepresenting academic or professional qualifications within or without the University; and nondisclosure or misrepresentation in filling out applications or other University records.

Sexual Misconduct and Gender Discrimination

In an effort to meet obligations under Title IX, UNM faculty, teaching assistants, and graduate assistants are considered "responsible employees." by the Department of Education (see page 15 -

<http://www2.ed.gov/about/offices/list/ocr/docs/qa-201404-title-ix.pdf>). This

designation requires that any report made to a faculty member, TA, or GA regarding sexual misconduct or gender discrimination must be reported to the Office of Equal Opportunity and the Title IX Coordinator. For more information on this policy, <https://policy.unm.edu/university-policies/2000/2740.html> or scan the QR Code above at right:



Title IX Policy